

27

2020 JUNE

DAY 179 - 187 WEEK 26

SATURDAY

06

June 2020

Wk	M	T	W	T	F	S	S
23	1	2	3	4	5	6	7
24	8	9	10	11	12	13	14
25	15	16	17	18	19	20	21
26	22	23	24	25	26	27	28
27	29	30					

APPOINTMENT / MEETING

Class VI

~~Date-28~~ Chapter-3 Matter Part 2
Subject Chemistry

Date-28.5.20

Discuss the properties of Solid, liquid and Gas.

Properties	Solid	liquid	Gas
(i) Intermolecular Space	Molecules are closely packed have negligible intermolecular space	Molecules are not closely packed have more intermolecular space	Molecules are at a greater distance from one other, free to move in all direction.
(ii) Shape	Have fixed shape	Have no fixed shape, take the shape of the container in which they are kept.	Have no fixed shape. Take the shape of the container in which they are filled.
(iii) Volume	Have fixed volume.	Have fixed volume.	Have no fixed volume.

Sunday 28

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Wk	M	T	W	T	F	S	S
27			1	2	3	4	5
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29	13	14	15	16	17	18	19
30	20	21	22	23	24	25	26
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APPOINTMENT / MEETING

Properties	Solid	Liquid	Gas
8 ① Fluidity	Do not flow	Flows from higher to lower level.	Flows in all directions
9			
10			
11 ② Effect of Pressure	Effect of pressure is very low	Effect of pressure is higher than on a solid	Effect of pressure is very high.
12 Compression			
1			
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27	29	30					

APPOINTMENT / MEETING

Q.2 Explain Solid state on the basis of characteristics of particles of matter.

Ans: In solids, the molecules are closely packed. There is a strong force of attraction between the molecules and spaces between them is very small (almost negligible). The molecules are therefore not free to move. They merely vibrate about their mean position. This makes solids hard and difficult to compress, giving them a fixed shape and size.



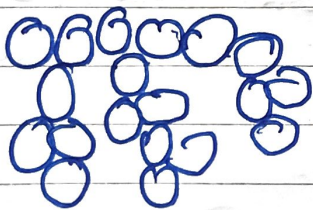
Molecules lie closely packed together, with great force of attraction between them.

Wk	M	T	W	T	F	S	S
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30	20	21	22	23	24	25	26
31	27	28	29	30	31		

APPOINTMENT / MEETING

Q.3 Explain liquid state on the basis of characteristics of molecules of matter.

Ans: In case of liquids, the molecules are not very closely packed. They do not attract each other strongly as the molecules of solids. Thus the intermolecular spaces are larger and the molecules are able to move more freely. This makes liquid flow and take the shape of the container into which it is poured. Thus liquid have a fixed volume but not definite shape of their own.



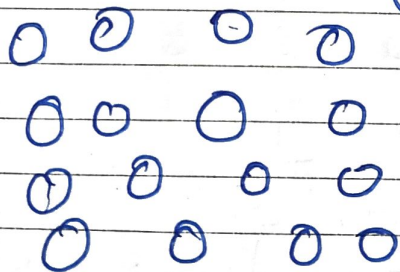
Molecules lie fairly less apart from each other, with little force of attraction between them.

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APPOINTMENT / MEETING

Q.4 Explain gaseous state on the basis of characteristics of molecules of matter.

Ans: In case of gases, the molecules hardly attract each other. They lie far apart from each other and the intermolecular spaces are therefore very large. The intermolecular force of attraction is so weak that the molecules have greater freedom of movement. As a result, gases have neither a fixed shape nor a fixed volume. They completely fill up the space available to them. They can be easily compressed and thus decrease the gap between them.



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Molecules lie very far from each other with very little force of attraction between them.

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31/36	31					1	2
32	3	4	5	6	7	8	9
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34	17	18	19	20	21	22	23
35	24	25	26	27	28	29	30

APPOINTMENT / MEETING

Q.5 What is Cohesive force?

Ans: The force of attraction between like particles or molecules is called cohesive force. Solids have strong cohesive force. Mercury, although a liquid, also shows strong cohesive force. Liquids have weaker cohesive force and gases have the weakest.

Q.6 Name two random motion of particles.

Ans: (i) Brownian movement (ii) Diffusion

Q.7 What is Brownian movement?

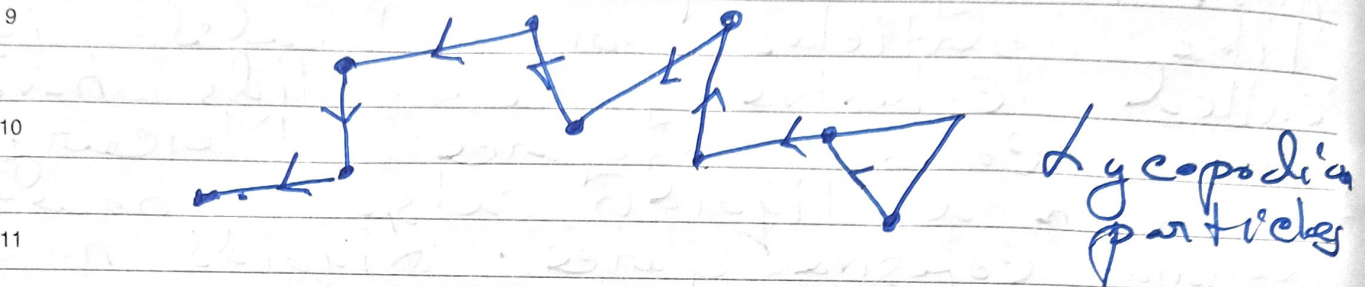
Ans: - The zigzag motion of particles suspended in a medium is called Brownian movement.

By: - Take a beaker half filled with water and suspend some Lycopodium (moss) powder in it. Look into the water through microscope. It will notice that the particles of Lycopodium are moving rapidly in zigzag (irregular) manner throughout

NOTES

APPOINTMENT / MEETING

8 The center. If the pebble is heated
the movement further increases



Q.8 What is Diffusion?

1

2

3

Ans: The phenomenon of intermingling of particles of one kind with another kind is called diffusion.

Q.9 Why are the gases - diffuse very fast?

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Ans: Gases diffuse very fast because the particles of gases are far enough apart which allows them to move freely and mix up easily.

Sunday 05

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31/36	31					1	2
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APPOINTMENT / MEETING

8 Q.10 what are miscible and immiscible liquids? Give eg.

9 Ans: liquids which mix with each other are called miscible liquids
10 Eg: Alcohol and water

11 liquids which do not mix with each other are called immiscible liquids
12 Eg: water and oil.

1 Q.11 why ~~are~~ solids cannot

2 Q.11 why do solids not diffuse?

3 Ans: Solids are very rigid, hence
4 either they do not diffuse under
5 normal conditions or diffuse
6 very little.

7 Q.12 what is the relationship between
8 temperature and diffusion?

9 Ans: When temperature increases, diffusion also increases.

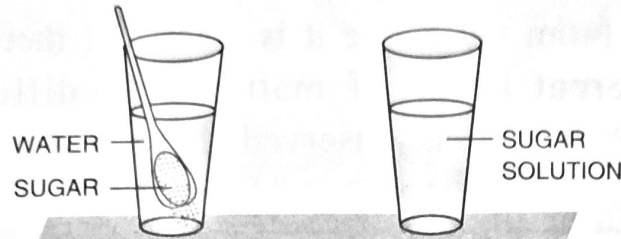
NOTES

Page-8

P. 55 ✓
20.12.20

Activity 2

Take half a glass of water. Add one teaspoon of sugar to it and stir. The sugar disappears but the level of water in the glass does not rise. It means the volume of water did not increase. But where did the sugar particles disappear ?

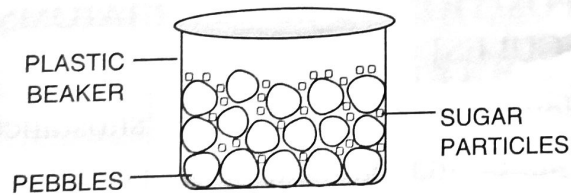


No change in the level of water taken

The sugar particles are adjusted between the water molecules. *This shows that there are intermolecular gaps in water.*

Activity 3

Take a plastic beaker and put some pebbles in it. Now add sugar or powdered table salt to it. You will observe that sugar particles get adjusted in the gaps between the pebbles.



Activity 4

Take a plastic beaker and put some glass balls in it. Now add sand to the beaker. You will observe that sand particles occupy the gaps between glass balls.

