Class - 6

Chapter - Matter

Q. Q.8. What do you mean by intermolecular forces?

Ans:- The force of attraction between the constituent particles is called the intermolecular force of attraction.

Q.9. What are the forces of cohesion and adhesion?

Ans:- The force of attraction between the particles of same substances is called the force of cohesion. and
The force of attraction between the particles of two different substances is called the force of adhesion.

Q.10 State three characteristics of molecules of matter which determine its solid, liquid and gaseous state.

Ans:- Three characteristics of molecules of matter which determine its solid, liquid and gaseous state are -

(i) inter-molecular space.
(ii) inter molecular force
(iii) movement of molecules.

Q.11 State the approximate spacing between two molecules of a matter.

Ans:- The approximate spacing between the two molecules of a matter ~~is~~ is nearly 10^{-9} m.

Q.12 How do solids, liquids and gases differ in their following properties:
(a) Size (b) Shape (c) Density?

Ans:-	Solid.	Liquid.	gas.
Size.	fixed.	Not fixed.	Not fixed.
Shape	definite	Acquires shape of the container.	Acquires the shape of container.
Density	high.	less than solids.	less dense.

Q.13. The molecules in a substance are in motion. What type of path do they follow?

Ans:- They follow a zig zag path.

Q.14 Describe a simple experiment to illustrate that molecules are not at rest but they constantly move.

Ans:- Take a beaker. Fill it partly with water. Add some lycopodium powder in the beaker containing water. Stir the contents of the beaker with a glass rod. Take out few drops of this suspension on a glass plate. Place the plate on the table and illuminate it with a table lamp. Observe the glass plate through a microscope. It is found that the fine particles of lycopodium powder move rapidly in a random manner and their path is zig zag.

Class - 6. Worksheet

1. Why do solids have a definite shape?

2. Why do liquids flow?

3. Why the gases are highly compressible?